

Characterization of Polymer–Solvent Interactions and Their Temperature Dependence Using Inverse Gas Chromatography

Minmin Tian and Petr Munk*

Department of Chemistry and Biochemistry and Center for Polymer Research, The University of Texas at Austin, Austin, Texas 78712

Inverse gas chromatography has been used to measure specific retention volumes and to derive a comprehensive set of interaction parameters for 9 polymers and 43 solvents at 6 temperatures from 60 to 110 °C. The specific retention volume, the Flory–Huggins interaction parameter χ_{12} , and the excess cohesive energy parameter B_{12} are presented for each polymer–solvent system as a function of temperature.

Introduction

Inverse gas chromatography (IGC), introduced about 20 years ago, is a rapid and reliable method for measuring polymer–solvent interactions (1–3). In a previous paper (4), we reported interaction parameters at 100 °C for 19 polymers and 43 solvents. We developed a multicomponent solubility parameter theory for polymer–solvent interactions (5) based on these data. As a continuation of our systematic investigations, in this paper we present interaction parameters gathered on 9 polymers with 43 solvents at six different temperatures. Analysis of these data will be presented in a separate paper.

During our IGC experiments, we have found that improvements in both the collection of experimental values and the data analysis (4, 6–9) are necessary to obtain more accurate thermodynamic values. The major sources of errors in common IGC experiments include (1) inaccurate determination of polymer mass on the IGC column by conventional solvent-extraction method, (2) poor constancy of the carrier-gas flow rate caused by room temperature fluctuations, and (3) improper analysis of the experimental results including neglection of the retention of the marker and of the retention of probes (the injected solvents) by the inert support. After elimination of these experimental errors and with proper data treatment, we estimate the accuracy of the specific retention volumes to be within 1%.

Theory

The change of the Gibbs free energy upon mixing, according to the Flory–Huggins theory, can be expressed as (10)

$$\Delta_{\text{mix}}G = RT[n_1 \ln \Phi_1 + n_2 \ln \Phi_2 + n_1 \Phi_1 \chi_{12}] \quad (1)$$

where Φ_1 and Φ_2 are the volume fractions of solvent and polymer, respectively, n_1 and n_2 are the numbers of moles of solvent and polymer, respectively, RT has its usual meaning, and χ_{12} is the Flory–Huggins interaction parameter.

Phenomenologically, the above equation has been considered as the definition of χ_{12} . All properties of a mixture can be expressed in terms of the variations of χ_{12} with composition, temperature, and pressure (11, 12). At the vanishing concentration of the probe, χ_{12} can be obtained

from the IGC experiment as

$$\chi_{12} = \ln(RTv_2/V_1P^{\circ}_1V_g) - 1 + V_1/M_2v_2 - (B_{11} - V_1)P^{\circ}_1/RT \quad (2)$$

where V_g is the specific retention volume of the probe, which is obtained directly from experiment, V_1 and v_2 are the solvent molar volume and polymer specific volume, respectively, P°_1 is the saturated vapor pressure of the probe, and B_{11} is the second virial coefficient of the probe in the vapor phase. The parameter of negative excess cohesive energy B_{12} is defined as

$$B_{12} = RT\chi_{12}/V_1 \quad (3)$$

The specific retention volume is determined as

$$V_g = j(t_p - t_m)F/w \quad (4)$$

where w is the mass of the polymer on the column and F is the flow rate of the carrier gas at the column temperature at the column outlet. t_p and t_m are the retention times of the probe and the marker, respectively. j is a correction factor to account for the compressibility of the carrier gas and is given by

$$j = (3/2)[(P_i/P_o)^2 - 1]/[(P_i/P_o)^3 - 1] \quad (5)$$

where P_i and P_o are the column inlet pressure and outlet pressure, respectively.

In many IGC experiments, column temperatures above 100 °C and often up to 200 °C are used. The saturated vapor pressure of some probes under these circumstances is very high; the nonideal gas virial term cannot be omitted and should be treated with care. Equation 2 becomes meaningless when the column temperature exceeds the critical temperature of the probe.

Experimental Section

Materials. Polyepichlorohydrin was purchased from Scientific Polymer Products Inc. and Goodrich Inc.; poly(ethylmethacrylate) was a gift from Dr. Fetter, Exxon Corp. All other polymers were purchased from Scientific Polymer Products Inc. Specific volumes of the polymers were either taken directly from the literature or calculated from the group contribution method. Polymer names, codes, and specific volumes are listed in Table 1.

All the probes were obtained from reputable suppliers. Only cyclopentane was of insufficient purity. It was distilled several times before use. Vapor pressures of the probes were obtained from their Antoine vapor pressure

* To whom correspondence should be addressed at the Department of Chemistry and Biochemistry.

Table 1. Listing of Polymers, Their Codes, and Physical Constants^a

polymer	code	T _g / °C	ϱ / (g/cm ³)	d ϱ /dT $\times 10^4$
poly(ethyl ethylene)	PEE	-4	0.872	6.69
polypropylene	PP1	-26	0.865	5.56
poly(dimethylsiloxane)	PDMS	-127	0.970	8.79 (22)
polycaprolactone	PCL	-60	1.095	8.64 (23)
polyepichlorohydrin	PECH	-22	1.360	10.4
poly(2-hydroxyethyl acrylate)	PHEA	-15	1.310	8.20
poly(methyl acrylate)	PMA	10	1.220	7.60 (24)
poly(ethyl acrylate)	PEA	-24	1.12	6.40
poly(<i>n</i> -butyl methacrylate)	PBMA	20	1.055	6.60 (25)

^a All ϱ and d ϱ /dT were calculated from ref 14, unless indicated otherwise.

Table 2. Listing of Probes and Their Critical Temperatures, Critical Volumes, and Critical Pressures

probe	code	T _c /K	V _c /(cm ³ /mol)	P _c /MPa
propane	NC3	370.0	197.5	4.261
butane	NC4	425.2	255.0	3.800
pentane	NC5	469.7	304.0	3.370
hexane	NC6	507.5	370.0	3.010
heptane	NC7	540.2	426.0	2.736
octane	NC8	568.8	492.0	2.488
nonane	NC9	594.6	548.0	2.288
decane	C10	617.7	603.0	2.120
undecane	C11	638.8	660.0	1.966
cyclopentane	CC5	511.9	259.8	4.517
cyclohexane	CC6	553.6	309.6	4.075
cycloheptane	CC7	604.4	353.0	3.860
cyclooctane	CC8	647.2	410.0	3.557
cyclohexene	CHX	560.5	292.0	
cyclohexadiene	CHD	561.5	276.4	
benzene	BNZ	562.2	260.3	4.898
toluene	TOL	594.0	320.0	4.236
ethylbenzene	EBZ	617.2	374.5	3.609
methyl chloride	CL1	416.2	139.0	6.679
methylene chloride	CL2	510.2	185.8	6.079
chloroform	CL3	536.4	240.7	5.472
carbon tetrachloride	CL4	556.4	275.9	4.560
butyl chloride	BCL	542.0	312.0	3.680
pentyl chloride	PCL	572.3	355.5	
chlorohexane	CLH	599.5	403.6	
chlorooctane	CLO	644.6	514.0	
1,1-dichloroethane	D11	523.0	240.0	5.070
1,2-dichloroethane	D12	566.0	225.0	5.370
methylchloroform	MCH	545.0	272.5	
trichloroethylene	TCE	572.0	256.0	5.050
chlorobenzene	CLB	632.4	308.1	4.520
acetone	ACT	508.1	209.0	4.700
methyl ethyl ketone	MEK	536.8	267.1	4.207
tetrahydrofuran	THF	540.2	224.0	4.770
dioxane	DOX	587.0	238.0	5.210
methyl acetate	MAC	506.8	228.0	4.690
ethyl acetate	EAC	523.2	286.0	3.830
propyl acetate	PAC	549.4	345.0	3.335
<i>n</i> -butyl acetate	NBA	579.0	400.0	3.140

coefficients at the proper temperature ranges. These coefficients together with other necessary constants were extracted from the Dreisbach compilation and other sources (14–18). Table 2 lists the probes and their codes, together with their critical temperatures, critical volumes, and critical pressures, which were used for the calculations of the gas virial coefficients.

Column Preparations. IGC columns were prepared from 0.64-cm-o.d. copper tubing 150 cm long, which was cleaned with methanol and left overnight in an oven at 110 °C. The support used was 60–80-mesh Chromosorb-W (acid washed and treated with dichlorodimethylsilane). The soaking method described previously was used for coating by the polymers. Polymer loadings were about 7% by mass.

Data Acquisition. The IGC experiments were performed on a modified Varian Aerograph 2100-40 gas

chromatograph equipped with a flame ionization detector (FID). Methane was used as a marker. The elution signal from the FID output was monitored by a Hewlett-Packard 3478A digital voltmeter and was processed in real time by a personal computer. Nitrogen was used as the carrier gas, and the flow rate was between 15 and 18 mL/min. The precision in the flow-rate measurement was ± 0.02 mL/min. Injected volumes were in the range of 0.02–0.04 μL and the retention volumes (after correction for the retention by the support; see below) were essentially independent of the injection sizes in a broad range. The temperature of the column was controlled within ± 0.1 °C.

During the data analysis, we have performed all the necessary corrections in the same way as we described previously (4, 5) except the correction for support retentions.

Retention by the Support. In our previous IGC experiments, we have found that the retention of the support is non-negligible in most cases and must be accounted for. The typical retention of the support for nonpolar probes accounts for 1–5% of the net retention volumes of the probe. For polar probes, it varied from 1% up to 20% depending on the amount of probe injected. Correction of the retention results for support adsorption can be performed by direct subtraction of the support retention from the apparent specific retention of the column, V_g^a, such that

$$V_g = V_g^a - V_n^s/w \quad (6)$$

where V_n^s is the retention of the probe on the support as obtained from an independent experiment on the uncoated support.

The support retentions depend on both temperature and the amount of probe injected. The concentration dependence of the support retentions varied with the chemical nature of the probe, being most pronounced for the strongly polar probes and negligible for alkanes and other hydrocarbons. Within the region of sample sizes practical to IGC experiments, all the concentration dependences of the probes were found to obey the relation

$$\ln V_n^s = a_1 + b_1 \ln A_p \quad (7)$$

where A_p, the peak area, is a measure of the probe sample size. a₁ and b₁ are temperature-dependent constants. In this study, we have found that the temperature dependence of support retentions obeys the Arrhenius relation

$$\ln V_n^s = a_2 + b_2/T \quad (8)$$

Combining eqs 7 and 8, we obtain

$$\ln V_n^s = a + b/T + c \ln A_p \quad (9)$$

Here, a₂, b₂, a, b, and c are probe-specific constants applicable for a given support and a given way of expressing the peak areas.

We have performed IGC experiments for each probe on the uncoated support at several temperatures. At each column temperature, we have injected different amounts of the probes. Thus, for each probe, we obtained a set of constants a, b, and c. The retention of the support for a given probe is then calculated from eq 9. At high temperatures, the retentions of lower carbon number hydrocarbons (propane, butane, etc.) by the support cannot be determined with good accuracy, because of extremely small retention volumes (<0.1 mL/g) of these probes. Nevertheless, we have found that for the same homologous group of compounds at low temperature, the retentions by the support are related to the carbon number with excellent

Table 3. Specific Retention Volume V_g of Various Probes on a PEE Column

$V_g/\text{cm}^3 \text{ g}^{-1}$						$V_g/\text{cm}^3 \text{ g}^{-1}$							
$t^\circ\text{C} =$	$t^\circ\text{C} =$	$t^\circ\text{C} =$	$t^\circ\text{C} =$	$t^\circ\text{C} =$	$t^\circ\text{C} =$	$t^\circ\text{C} =$	$t^\circ\text{C} =$	$t^\circ\text{C} =$	$t^\circ\text{C} =$	$t^\circ\text{C} =$	$t^\circ\text{C} =$		
60.0	70.0	80.0	90.0	100.0	110.0	60.0	70.0	80.0	90.0	100.0	110.0		
NC3	4.90	4.06	3.71	3.50		BCL	117.15	88.15	67.89	53.10	42.06	34.23	
NC4	13.47	10.80	9.53	8.19		PCL	294.14	211.80	155.80	116.63	88.72	69.63	
NC5	34.52	26.90	21.98	17.73	14.85	12.34	CLH	722.32	495.98	349.11	250.43	182.93	138.55
NC6	86.87	65.17	50.39	39.59	31.34	25.25	CLO	4214.19	2644.18	1723.32	1138.36	761.30	535.33
NC7	214.54	153.73	114.56	85.77	65.18	50.60	D11	56.27	44.22	35.32	28.72	23.47	19.69
NC8	522.71	357.23	255.39	182.84	133.68	100.23	D12	102.50	78.50	61.23	48.68	38.70	31.98
NC9	1263.64	825.41	565.08	386.77	272.42	196.55	MCH	116.98	89.63	68.45	54.93	43.55	35.83
C10	2991.01	1906.43	1241.73	815.31	550.53	384.21	TCE	207.44	153.32	114.18	88.97	68.91	55.28
C11	4360.78	2697.40	1755.65	1123.99	752.73	CLB	696.61	489.22	348.60	258.08	191.65	147.93	
CC5	71.88	55.69	44.62	35.27	28.20	23.70	ACT	15.96	13.33	11.70	10.05	8.40	7.60
CC6	178.66	133.29	102.71	78.74	61.52	49.24	MEK	48.38	37.94	30.02	24.60	19.84	16.89
CC7	602.92	424.55	306.93	226.73	168.80	130.52	THF	89.66	68.64	53.26	42.68	33.86	28.19
CC8	1780.02	1189.13	810.83	577.37	412.66	307.44	DOX	152.39	112.39	85.60	67.29	51.83	42.77
CHX	193.82	143.93	109.24	84.18	65.75	52.79	MAC	25.21	20.15	16.49	13.60	11.11	9.70
CHD	163.21	122.06	93.63		57.11	46.42	EAC	54.88	41.89	32.83	26.37	20.82	17.51
BNZ	140.70	106.28	81.97	64.31	50.87	41.47	PAC	137.34	100.48	74.94	57.81	44.24	35.67
TOL	359.94	258.93	191.82	143.22	108.80	85.62	NBA	341.47	238.14	169.61	125.47	92.56	72.06
EBZ	804.68	556.78	395.33	284.79	209.45	159.25	EOH	9.27	7.45	7.58	6.31	4.65	4.93
CL1	6.84	5.85	5.73	4.74	4.56	3.84	POH	30.40	23.75	19.99	16.28	13.08	11.64
CL2	34.24	27.43	22.90	18.62	15.70	13.55	BOH	82.11	61.89	48.22	38.12	29.65	25.26
CL3	84.81	65.57	52.14	41.36	33.34	27.80	AOH	201.14	152.17	112.39	85.47	64.61	52.02
CL4	155.48	117.57	90.87	70.51	55.66	45.33							

Table 4. Specific Retention Volume V_g of Various Probes on a PP1 Column

$V_g/\text{cm}^3 \text{ g}^{-1}$					$V_g/\text{cm}^3 \text{ g}^{-1}$						
$t^\circ\text{C} =$	$t^\circ\text{C} =$	$t^\circ\text{C} =$	$t^\circ\text{C} =$	$t^\circ\text{C} =$	$t^\circ\text{C} =$	$t^\circ\text{C} =$	$t^\circ\text{C} =$	$t^\circ\text{C} =$	$t^\circ\text{C} =$		
70.0	80.0	90.0	100.0	110.0	70.0	80.0	90.0	100.0	110.0		
NC3	3.95	3.62	2.65		BCL	84.52	62.31	50.38	41.62	32.69	
NC4	9.97	8.91	6.71		PCL	203.72	152.87	112.50	88.36		
NC5	25.66	20.61	16.78	11.36	CLH	486.20	332.68	244.46	184.30	142.21	
NC6	63.52	49.32	37.75	29.87	25.13	CLO	2661.89	1732.12	1136.18	778.78	571.05
NC7	153.08	113.55	83.60	63.84		D11	41.28	33.59	26.61	22.91	18.43
NC8	360.45	254.02	181.49	134.04	105.87	D12	74.92	59.20	46.07	38.46	30.35
NC9	764.17	571.99	388.52	276.58	211.79	MCH	83.30	65.09	50.93	42.04	34.64
C10	1956.05	1281.77	826.07	568.46	436.03	TCE	145.72	108.79	84.69	67.77	54.31
C11	4524.80	2913.40	1750.84	1158.50	828.20	CLB	470.97	328.04	246.32	189.82	151.30
CC5	52.90	41.79	32.90	26.48	22.12	ACT	11.63	11.64	9.13	8.71	7.55
CC6	127.45	96.81		58.80	47.87	MEK	36.35	29.02	23.44	20.26	16.18
CC7	408.05	295.72	214.92	162.87	130.33	THF	64.50	50.94	40.34	33.86	26.70
CC8	1121.03	764.13	552.59	402.61	310.24	DOX	106.98	82.82	63.83	52.70	41.40
CHX	136.82	103.78	79.30	63.31	51.03	MAC	17.98	16.51	12.49	11.60	9.80
CHD	115.63	88.41	68.08	55.16	44.46	EAC	40.10	31.97	24.81	21.34	16.29
BNZ	101.55	77.20	60.31	48.94	40.10	PAC	97.64	73.75	55.60	45.05	37.11
TOL	249.44	185.68	138.91	107.89	86.48	NBA	234.53	171.51	122.56	95.17	73.36
EBZ	542.19	387.19		212.96	161.92	EOH	8.22	7.83	5.25	4.99	4.74
CL1	6.02	5.23	4.64	4.44	3.80	POH	21.41	19.69	15.27	13.14	11.98
CL2	25.17	21.88	17.70	15.25	11.06	BOH	58.71	45.60	36.52	29.55	24.46
CL3	60.32	47.37	38.93	32.01	26.89	AOH	147.14	108.84	82.43	65.62	52.15
CL4	105.72	82.97	64.67	52.78	43.21						

accuracy as

$$\ln V_n^s = d + eN \quad (10)$$

where N is the number of carbons in the molecule and d and e are constants. We can therefore obtain the support retentions of low molecular weight hydrocarbons by extrapolating V_n^s vs. carbon number according to eq 10. By controlling the injected volume, we have ascertained that the retention of the support was never greater than 5% of the net retention volume of the probe on coated column. In most cases, the retention of the support was within 2% of the corrected retention volume of the probe.

We have also ascertained that our values of V_g were not adversely affected by slow diffusion of probes through the polymer layer (19–21). We have demanded that the height of the equivalent theoretical plate (HETP) (which is the measure of diffusion of the probes on the column and varies from probe to probe) be the same for a given probe when measured on coated and uncoated columns. We have found that all our measurements satisfied this criterion except

for PHEA and PMA at low temperatures. Even for these measurements, the estimated errors in V_g , χ_{12} , and B_{12} were rather small. Nevertheless, we are excluding these values when analyzing the temperature dependence of the interaction parameters.

Results

The experimental values of the specific retention volume V_g are collected in Tables 3–11. All thermodynamic information about polymer–solvent interactions, their temperature dependence, equation-of-state effects, Henry's constants for each probe, etc., can be directly derived from the V_g values. We have listed the interaction parameters χ_{12} defined by eq 1 in Tables 12–20 and excess free energy of mixing B_{12} in Tables 21–29. We have excluded the χ_{12} and B_{12} values of propane and butane at 100 and 110 °C since the column temperatures are near or exceed the critical temperatures of the probes. We have also excluded the interaction parameters of propane, butane, pentane, and hexane on poly(2-hydroxyethyl acrylate), because of

Table 10 (Continued)

	V _g /(cm ³ g ⁻¹)					V _g /(cm ³ g ⁻¹)					
	t/°C = 70.0	t/°C = 80.0	t/°C = 90.0	t/°C = 100.0	t/°C = 110.0		t/°C = 70.0	t/°C = 80.0	t/°C = 90.0	t/°C = 100.0	t/°C = 110.0
CC8	288.58	199.51	148.39	110.39	86.11	TCE	150.05	110.09	83.08	63.32	48.73
BNZ	118.56	88.57	68.39	52.21	41.61	CLB	613.85	425.56	305.80	218.26	162.85
TOL	240.58	179.06	129.01	94.78		ACT	49.51	37.92		23.19	20.64
EBZ	448.99	311.00	225.44	163.60	129.32	MEK	95.20	71.21	53.69	41.95	39.51
CL2	70.07	53.61	41.98	32.72		THF	79.26	60.84	46.50	36.97	
CL3	143.18	104.40	78.89	59.69	47.02	DOX	255.13	183.19	132.99	99.85	
CL4	80.14	60.73	47.33	37.01	30.13	MAC	51.45	35.89	29.15	24.18	19.50
BCL	70.33	53.93	42.35	32.97	26.86	EAC	83.04	61.42	45.86	36.23	28.42
PCL	145.46	107.86	81.07	61.11	49.37	PAC	162.15	115.91	84.05	63.81	48.97
CLH		209.33	152.38	112.48	85.38	NBA	329.30	221.51	158.08	115.80	86.76
CLO			533.15	366.12	260.48	EOH	66.81	49.80	38.03	29.09	23.03
D11	72.93	55.40	43.26	33.87		POH	147.75	102.18	75.64	56.29	43.45
D12	194.64	141.95	106.26	79.64	61.63	BOH	318.91	215.39	150.65	109.17	81.83
MCH	84.92	64.34	50.16	39.52	31.14						

Table 11. Specific Retention Volume V_g of Various Probes on PBMA

	V _g /(cm ³ g ⁻¹)					V _g /(cm ³ g ⁻¹)					
	t/°C = 70.0	t/°C = 80.0	t/°C = 90.0	t/°C = 100.0	t/°C = 110.0		t/°C = 70.0	t/°C = 80.0	t/°C = 90.0	t/°C = 100.0	t/°C = 110.0
NC3	2.62	2.23	1.85			BCL	90.97	68.71	56.00	43.04	32.75
NC4	5.73	4.86	4.20			PCL	208.07	151.56	119.59	89.29	40.81
NC5	13.09	9.82	9.14	7.46	6.09	CLH	463.88	324.20	230.48	178.41	127.02
NC6	30.43	23.78	19.60	15.70	12.52	CLO	2272.02	1463.36	953.25	696.52	460.69
NC7	69.68	52.61	40.89	31.78	25.06	D11	72.11	54.81	44.80	34.71	26.83
NC8	153.63	111.60	83.38	63.51	48.62	D12	166.13	122.08	96.04	73.01	54.54
NC9	338.45	236.01	173.33	124.82	92.34	MCH	96.17	73.67	56.83	46.86	36.02
C10	744.65	496.89	351.72	243.81	173.59	TCE	180.38	131.04	97.58	77.91	58.22
C11	1628.94	1044.87	709.32	474.58	326.15	CLB	700.02	489.95	346.67	266.64	190.25
CC5	29.05	23.57	19.43	15.90	12.72	ACT	34.49	26.91	21.28	17.91	14.11
CC6	63.78	50.75	41.38	33.13	26.72	MEK	78.83	59.28	45.47	36.79	28.65
CC7	201.89	152.26	119.31	90.57	69.01	THF	80.13	61.03	44.96	38.64	29.91
CC8	559.66	405.10	297.11	220.50	164.55	DOX	202.27	145.39	108.58	86.48	64.85
CHX	90.00	69.68	55.90	44.57	34.74	MAC	37.99	29.18	22.65	18.82	14.60
CHD	102.28	78.54	62.60	49.73	37.61	EAC	69.84	51.75	39.15	31.92	24.38
BNZ	128.34	96.91	76.09	57.62	44.99	PAC	152.81	108.79	80.28	63.57	47.05
TOL	290.08	209.76	159.72	116.76	87.68	NBA	344.79	235.40	167.02	128.80	92.53
EBZ	594.51	413.53	290.80	216.46	157.11	EOH	44.93	33.23	24.65	20.18	15.36
CL1	8.13	6.95	5.83	5.39	4.68	POH	112.13	79.13	57.87	45.69	34.07
CL2	59.16	45.69	36.82	28.92	22.28	BOH	270.86	182.49	128.92	99.43	70.77
CL3	139.65	102.84	79.78	61.55	45.66	AOH	633.04	410.05	279.68	208.21	142.25
CL4	97.24	74.44	61.08	47.66	36.77						

Table 12. Flory–Huggins Interaction Parameters χ₁₂ for PEE

	χ ₁₂ (unitless)						χ ₁₂ (unitless)					
	t/°C = 60.0	t/°C = 70.0	t/°C = 80.0	t/°C = 90.0	t/°C = 100.0	t/°C = 110.0		t/°C = 60.0	t/°C = 70.0	t/°C = 80.0	t/°C = 90.0	t/°C = 100.0
NC3	0.454	0.494	0.441	0.366			BCL	0.548	0.525	0.508	0.485	0.472
NC4	0.396	0.416	0.361	0.340			PCL	0.465	0.436	0.420	0.395	0.380
NC5	0.349	0.360	0.338	0.345	0.324	0.326	CLH	0.425	0.401	0.392	0.375	0.369
NC6	0.302	0.304	0.293	0.286	0.285	0.287	CLO	0.402	0.373	0.354	0.340	0.336
NC7	0.271	0.272	0.255	0.255	0.257	0.260	D11	0.805	0.762	0.731	0.691	0.662
NC8	0.252	0.253	0.234	0.238	0.241	0.243	D12	1.122	1.071	1.032	0.983	0.953
NC9	0.243	0.214	0.220	0.228	0.228	0.233	MCH	0.416	0.381	0.369	0.333	0.319
C10	0.236	0.210	0.212	0.218	0.221	0.223	TCE	0.405	0.387	0.383	0.362	0.356
C11	0.218	0.203	0.190	0.200	0.217		CLB	0.603	0.572	0.552	0.529	0.514
CC5	0.196	0.193	0.173	0.184	0.196	0.177	ACT	2.156	2.043	1.899	1.805	1.741
CC6	0.155	0.147	0.129	0.136	0.137	0.135	MEK	1.628	1.552	1.488	1.418	1.369
CC7	0.078	0.074	0.068	0.064	0.068	0.060	THF	0.647	0.614	0.587	0.557	0.541
CC8	0.034	0.033	0.032	0.032	0.037	0.029	DOX	1.351	1.299	1.237	1.180	1.148
CHX	0.211	0.203	0.194	0.192	0.189	0.181	MAC	1.646	1.574	1.497	1.442	1.403
CHD	0.365	0.352	0.336	0.323	0.306		EAC	1.346	1.293	1.234	1.183	1.154
BNZ	0.579	0.550	0.524	0.499	0.484	0.460	PAC	1.153	1.097	1.047	0.999	0.966
TOL	0.490	0.468	0.443	0.431	0.422	0.401	NBA	1.004	0.958	0.917	0.880	0.851
EBZ	0.458	0.434	0.412	0.400	0.391	0.373	EOH	3.813	3.628	3.231	3.072	2.716
CL1	0.897	0.863	0.709	0.736	0.621	0.662	POH	3.213	3.014	2.770	2.602	2.457
CL2	0.996	0.953	0.888	0.866	0.825	0.777	BOH	2.964	2.748	2.534	2.358	2.209
CL3	0.638	0.626	0.606	0.607	0.609	0.594	AOH	2.720	2.470	2.284	2.124	1.984
CL4	0.268	0.251	0.234	0.230	0.226	0.211						

Table 21 (Continued)

B ₁₂ /J/cm ³)						B ₁₂ /J/cm ³)							
t/°C =	t/°C =	t/°C =	t/°C =	t/°C =	t/°C =	t/°C =	t/°C =	t/°C =	t/°C =	t/°C =	t/°C =		
60.0	70.0	80.0	90.0	100.0	110.0	60.0	70.0	80.0	90.0	100.0	110.0		
NC7	4.86	4.96	4.72	4.79	4.88	5.00	CLH	8.21	7.90	7.85	7.64	7.64	7.40
NC8	4.09	4.19	3.93	4.07	4.16	4.26	CLO	6.29	5.95	5.75	5.63	5.66	5.64
NC9	3.60	3.64	3.38	3.56	3.61	3.75	D11	25.04	24.05	23.38	22.35	21.65	20.88
C10	3.22	2.92	2.99	3.13	3.22	3.30	D12	37.52	36.40	35.63	34.43	33.84	32.78
C11	2.80	2.66	2.53	2.71	2.99	MCH	10.99	10.21	10.02	9.18	8.88	8.24	
CC5	5.47	5.45	4.96	5.33	5.73	5.20	TCE	11.88	11.56	11.61	11.14	11.10	10.81
CC6	3.77	3.63	3.26	3.46	3.54	3.52	CLB	15.75	15.23	14.98	14.60	14.42	14.07
CC7	1.71	1.64	1.54	1.49	1.59	1.41	ACT	76.63	73.60	69.23	66.51	64.74	60.95
CC8	0.68	0.67	0.67	0.67	0.78	0.63	MEK	47.82	46.30	45.02	43.49	42.48	40.61
CHX	5.49	5.36	5.21	5.22	5.21	5.05	THF	20.74	19.90	19.22	18.41	18.00	16.92
CHD	10.14	9.94	9.62	9.50	9.10	DOX	43.46	42.63	41.40	40.21	39.83	38.17	
BNZ	17.15	16.58	16.02	15.47	15.20	14.60	MAC	54.11	52.43	50.47	49.15	48.23	45.89
TOL	12.21	11.89	11.43	11.30	11.23	10.82	EAC	36.02	35.09	33.94	32.92	32.46	31.04
EBZ	9.94	9.60	9.29	9.17	9.09	8.82	PAC	26.37	25.51	24.69	23.90	23.38	22.43
CL1	40.17	38.52	31.41	32.20	26.57	27.32	NBA	20.12	19.53	19.00	18.51	18.16	17.48
CL2	40.58	39.34	37.12	36.60	35.15	33.35	EOH	172.81	167.23	151.28	145.91	147.56	132.27
CL3	20.82	20.73	20.34	20.62	20.91	20.60	POH	113.97	108.80	101.66	96.99	92.88	86.25
CL4	7.30	6.97	6.58	6.56	6.54	6.16	BOH	86.13	81.32	76.32	72.19	68.68	63.63
BCL	13.81	13.47	13.22	12.80	12.61	12.22	AOH	66.92	61.91	58.25	55.09	52.24	48.95
PCL	10.17	9.70	9.50	9.09	8.86	8.40							

Table 22. Parameters B_{12} for PP1

$B_{12}/(\text{J/cm}^3)$					$B_{12}/(\text{J/cm}^3)$					
$t/\text{^{\circ}C} =$	$t/\text{^{\circ}C} =$	$t/\text{^{\circ}C} =$	$t/\text{^{\circ}C} =$	$t/\text{^{\circ}C} =$	$t/\text{^{\circ}C} =$	$t/\text{^{\circ}C} =$	$t/\text{^{\circ}C} =$	$t/\text{^{\circ}C} =$	$t/\text{^{\circ}C} =$	
70.0	80.0	90.0	100.0	110.0	70.0	80.0	90.0	100.0	110.0	
NC3	14.09	12.16	16.93	18.60	BCL	14.69	15.45	14.19	12.82	13.15
NC4	12.83	10.77	13.60	14.69	PCL	11.10	10.40	10.24	9.62	
NC5	9.43	9.20	9.27		CLH	8.45	8.62	8.08	7.40	6.44
NC6	6.85	6.41	6.95	6.92	CLO	6.05	5.66	5.61	5.52	4.26
NC7	5.20	4.85	5.29	5.21	D11	26.44	25.00	25.01	22.80	22.94
NC8	4.21	3.96	4.20	4.05	D12	38.23	36.79	36.66	34.53	34.64
NC9	4.94	3.13	3.53	3.31	MCH	12.36	11.65	11.63	10.21	9.44
C10	2.67	2.47	2.97	2.69	TCE	13.29	13.30	12.96	12.05	11.39
C11	2.56	1.73	2.57	2.32	CLB	16.46	16.68	15.77	15.10	13.41
CC5	7.12	6.91	7.18	7.49	ACT	78.84	69.66	71.35	63.85	60.72
CC6	5.00	4.82		4.63	MEK	47.80	46.34	45.77	42.22	42.73
CC7	2.78	2.49	2.66	2.34	THF	22.15	21.00	20.12	18.41	18.27
CC8	2.04	2.05	1.54	1.22	DOX	44.50	42.86	42.57	39.75	39.78
CHX	6.96	6.70	6.42	6.16	MAC	56.54	50.66	48.18	47.15	45.07
CHD	11.75	11.38	11.57	10.43	EAC	36.47	34.94	34.92	32.13	33.23
BNZ	18.26	17.98	17.31	16.32	PAC	26.36	25.33	24.95	23.30	21.06
TOL	13.00	12.40	12.17	11.37	NBA	20.01	18.95	19.09	17.91	17.01
EBZ	10.33	9.79		8.67	EOH	163.19	150.13	156.18	143.76	133.44
CL1	37.53	35.51	34.29	27.61	POH	112.98	102.52	99.74	93.35	84.44
CL2	43.10	39.21	38.77	36.28	BOH	83.16	78.30	73.61	69.37	64.50
CL3	21.39	20.89	20.28	18.25	AOH	62.99	59.33	56.11	52.34	48.72
CL4	10.07	9.26	9.21	7.99						

Table 23. Parameters B_{12} for PDMS

Table 26 (Continued)

B ₁₂ /J/cm ³						B ₁₂ /J/cm ³					
t/°C =	t/°C =	t/°C =	t/°C =	t/°C =	t/°C =	t/°C =	t/°C =	t/°C =	t/°C =	t/°C =	t/°C =
60.0	70.0	80.0	90.0	100.0	110.0	60.0	70.0	80.0	90.0	100.0	110.0
CLH	71.16	67.56	65.14	64.11	63.29	62.34	DOX	24.35	23.13	22.91	22.79
CLO	64.74	63.76	63.04	62.86	62.60	62.36	MAC	42.02	40.25	39.82	39.17
D11	63.06	56.83	53.69	50.98	50.48	49.12	EAC	44.43	42.41	41.58	40.59
D12	42.26	40.67	40.27	39.95	40.55	40.63	PAC	46.02	43.87	42.98	42.10
MCH		70.87	65.25	59.98	57.68	55.87	NBA	45.70	44.24	43.54	42.91
TCE	64.05	60.97	59.63	58.45	58.83	58.38	EOH	36.80	35.95	35.07	34.56
CLB	51.47	50.09	49.58	49.21	49.74	49.61	POH	34.93	33.48	32.64	31.49
ACT	34.87	32.38	31.25	30.24	30.02	29.44	BOH	35.33	34.12	32.82	32.00
MEK	36.99	35.04	34.25	33.45	33.74	32.98	AOH	33.36	32.72	32.02	31.49
THF	36.91	33.79	32.40	31.23	31.92	30.61					

Table 27. Parameters B₁₂ for PMA

B ₁₂ /J/cm ³						B ₁₂ /J/cm ³					
t/°C =	t/°C =	t/°C =	t/°C =	t/°C =	t/°C =	t/°C =	t/°C =	t/°C =	t/°C =	t/°C =	t/°C =
60.0	70.0	80.0	90.0	100.0	110.0	60.0	70.0	80.0	90.0	100.0	110.0
NC3	65.47	62.31	46.84	43.87		BCL	24.28	22.70	21.66	20.69	20.33
NC4	60.25	56.38	47.14	43.92		PCL	23.74	22.38	21.38	20.33	19.83
NC5	61.40	53.10	46.73	45.21	42.66	42.10	CLH	24.02	22.89	22.31	21.47
NC6	54.61	48.80	44.88	42.21	40.93	39.77	CLO	25.22	24.42	23.88	23.15
NC7	50.40	45.87	42.95	40.65	39.58	38.06	D11	9.08	7.53	7.36	6.82
NC8	47.04	43.71	41.49	39.67	38.80	37.51	D12	-0.86	-0.76	-0.05	-0.04
NC9	44.51	42.14	40.53	38.93	37.97	36.95	MCH	19.28	15.35	13.84	12.45
C10	42.31	40.59	39.53	38.17	37.24	36.02	TCE	14.20	13.56	13.83	13.35
C11	41.22	40.05	38.93	37.70	36.77	35.57	CLB	10.38	9.49	8.90	8.63
CC5	59.57	51.13	47.15	43.12	41.78	41.43	ACT	17.14	16.43	16.24	15.03
CC6	56.49	48.59	44.04	40.91	39.28	38.16	MEK	13.49	13.05	12.96	12.26
CC7	48.37	43.17	39.97	37.58	36.20	35.09	THF	13.61	12.72	11.32	10.97
CC8	43.13	39.62	37.21	35.48	34.43	33.41	DOX	6.61	6.35	6.52	6.26
CHX	41.11	36.75	34.30	32.11	30.94	30.32	MAC	13.02	13.32	13.00	12.56
CHD	28.42	25.93	24.56	23.20	22.67	22.54	EAC	12.59	12.30	11.95	11.89
BNZ	13.96	12.74	12.10	11.33	11.36	11.41	PAC	13.09	12.56	12.16	11.79
TOL	15.57	14.87	14.39	13.79	13.77	13.81	NBA	13.18	12.78	12.41	12.14
EBZ	17.38	16.78	16.30	15.69	15.58	15.55	EOH	58.91	54.95	53.02	49.70
CL1	17.34	15.21	13.46	11.08	9.95	9.50	POH	39.49	37.69	35.41	32.77
CL2	-5.59	-4.97	-4.57	-4.44	-3.64	-2.42	BOH	33.56	30.70	28.40	25.95
CL3	-7.26	-6.79	-5.88	-5.29	-3.81	-2.58	AOH	27.57	25.03	22.85	21.03
CL4	26.91	23.05	21.17	19.74	19.40	19.23					

Table 28. Parameters B₁₂ for PEA

B ₁₂ /J/cm ³						B ₁₂ /J/cm ³					
t/°C =	t/°C =	t/°C =	t/°C =	t/°C =	t/°C =	t/°C =	t/°C =	t/°C =	t/°C =	t/°C =	t/°C =
70.0	80.0	90.0	100.0	110.0		70.0	80.0	90.0	100.0	110.0	
NC3	32.34	25.67	21.46			PCL	12.84	12.49	12.23	12.25	11.34
NC4	33.51	33.05	31.39			CLH		12.95	12.61	12.32	12.14
NC5	31.88	30.25	29.83	28.55	28.75	CLO			13.87	13.74	13.56
NC6	30.07		28.90		27.38	D11	0.39	0.87	0.88	1.38	
NC7	28.91	28.07	27.21	26.51	25.82	D12	-2.87	-2.10	-1.71	-0.67	0.16
NC8	27.95	27.40	26.71	26.21	25.15	MCH	5.36	5.53	5.29	5.32	5.88
NC9	27.41	26.95	26.19	25.83	24.83	TCE	4.86	5.43	5.58	6.08	6.92
C10	26.72	26.12	25.61	25.33	24.35	CLB	2.51	2.92	2.85	3.61	3.98
C11	26.58	25.94	25.31	25.04	24.10	ACT	18.25	18.47		18.91	15.41
CC5	29.12	28.16	27.16	27.34	26.98	MEK	11.94	12.08	12.36	12.18	
CC6	28.46	27.71	26.50	26.11	25.49	THF	6.97	6.64	6.83	6.37	
CC7	26.55	25.93	24.78	24.19	23.26	DOX	7.84	8.15	8.65	8.85	
CC8	24.25	24.49	23.73	23.27	22.53	MAC	13.39	16.57	15.07	13.51	13.66
BNZ	5.51	5.73	5.46	6.00	5.99	EAC	9.90	10.17	10.56	10.06	10.30
TOL	7.34	6.69	7.38	8.00		PAC	8.63	8.71	8.82	8.53	8.49
EBZ	8.76	9.00	8.70	8.82	7.59	NBA	8.02	8.47	8.23	8.10	8.00
CL2	-9.92	-9.09	-8.59	-7.35		EOH	55.80	52.83	49.69	48.92	46.76
CL3	-15.86	-14.32	-13.40	-12.06	-11.18	POH	34.26	33.07	30.33	28.59	26.42
CL4	10.67	10.91	10.80	10.95	10.67	BOH	25.72	24.00	22.19	20.46	18.59
BCL	12.73	12.46	12.00	12.10	11.69						

Table 29. Parameters B₁₂ for PBMA

B ₁₂ /J/cm ³						B ₁₂ /J/cm ³					
t/°C =	t/°C =	t/°C =	t/°C =	t/°C =	t/°C =	t/°C =	t/°C =	t/°C =	t/°C =	t/°C =	t/°C =
70.0	80.0	90.0	100.0	110.0		70.0	80.0	90.0	100.0	110.0	
NC3	19.81	20.08	20.96			NC7	15.79	15.51	14.89	14.72	14.60
NC4	21.72	21.24	20.39			NC8	14.86	14.57	14.00	13.56	13.37
NC5	20.07	21.78	18.63	18.89	19.36	NC9	14.11	13.82	12.92	12.82	12.64
NC6	17.58	17.42	16.34	16.26	16.58	C10	13.20	12.94	12.28	12.17	12.05

of the experiments were generally not sufficiently close to each other: we will therefore not compare our data to the literature. In order to get some feeling about the effect of the experimental arrangement on IGC data, we report in Table 31 the χ_{12} values measured by three different researchers in our laboratory, using three different gas chromatographs and two samples of the polymer (PECH) obtained from two different manufacturers. Except for a few probes with a short retention time (propane, butane, methyl chloride—all gases at room temperature with V_g values less than 4 mL/g), and ethanol with a large χ_{12} value, the average differences of χ_{12} are less than 0.01. The reliability of our results can also be determined from the inspection of the data themselves, *i.e.*, from judging the scatter of the χ_{12} values from a presumably smooth dependence on the temperature, or on the number of carbons in a family of related probes interacting with any given polymer or, conversely, for a given probe interacting with a family of similar polymers. Definite correlations are observed. We concluded that our IGC experimental data are reproducible and reliable.

A straightforward thermodynamic calculation links the temperature dependence of V_g to the molar enthalpy of sorption of the probe by the polymer, $\Delta_{\text{sorp}}H$, as

$$\frac{d \ln(V_g/T)/d(1/T)}{d} = -\Delta_{\text{sorp}}H/R \quad (11)$$

Table 30 lists values of $-\Delta_{\text{sorp}}H$ for nine polymers at six different temperatures. In the last column; the heat of vaporization of the probe at 85 °C is also listed. When calculating $\Delta_{\text{sorp}}H$, we have excluded all V_g values smaller than 10 mL/g. We have also excluded the V_g values obtained on the poly(methyl acrylate) column at 60 °C and on the poly(2-hydroxyethyl acrylate) column at 60 and 70 °C. These measurements were slightly influenced by the slow diffusion of the probes in the polymer phase. While we have estimated that the error caused by the diffusion influences the V_g , χ_{12} , and B_{12} values only to an acceptable degree, we have found that it distorts the V_g vs 1/T dependence beyond limits acceptable for calculation of $\Delta_{\text{sorp}}H$.

Equation 11 fitted V_g data quite well. The standard deviation of V_g values was less than 3%, and the absolute error of $\Delta_{\text{sorp}}H$ was less than 0.8 kJ/mol; this compares with the typical value of $\Delta_{\text{sorp}}H$, -25 kJ/mol. It is apparent that for solvents good for a given polymer, $-\Delta_{\text{sorp}}H$ is close to $\Delta_{\text{vap}}H$, while for poor solvents, $-\Delta_{\text{sorp}}H$ is less than $\Delta_{\text{vap}}H$.

It is seen from experimental values listed in the tables that most the solvent–polymer interaction parameters decrease with increasing temperature. Poly(2-hydroxyethyl acrylate) has the largest χ_{12} values for most of our probes, indicating that it has poor interactions with all of our probes (except alcohols). Chloro compounds with slightly acidic character, such as methylene chloride, chloroform, 1,1-dichloroethane, 1,2-dichloroethane, etc., have negative interaction parameters with polycaprolactone, poly(methyl acrylate), and poly(*n*-butyl methacrylate), *i.e.*, with polymers that contain a carboxyl group in the backbone chain or in side groups. This shows strong interaction between acidic moieties and -COO groups. Aromatic compounds, as well as organic ethers, show

moderately strong interactions with polyepichlorohydrin. We will present a detailed thermodynamic analysis in another paper.

Conclusions

(1) IGC is a quick and reliable method for obtaining the polymer–solvent interaction parameters as well as other thermodynamic properties at the vanishing concentration of solvents. The interaction parameters of 43 common solvents and 9 polymers have been determined in this study. Each polymer–solvent interaction parameter was determined at six different temperatures from 60 to 110 °C.

(2) A number of experimental and analysis refinements were implemented in improving the precision and reproducibility of the thermodynamic data.

(3) Extreme care should be taken in the data treatment when the column temperature nears the critical temperature of a probe; otherwise invalid thermodynamic results will be obtained.

Literature Cited

- (1) Smidsrod, O.; Guillet, J. E. *Macromolecules* **1969**, *2*, 273.
- (2) Patterson, D.; Tewari, Y. B.; Schreiber, H. P.; Guillet, J. E. *Macromolecules* **1971**, *4*, 356.
- (3) Walsh, D. J.; McKeown, J. G. *Polymer* **1980**, *21*, 1335.
- (4) Munk, P.; Hattam, P.; Du, Q.; Abdel-Azim, A.-A. *J. Appl. Polym. Sci., Appl. Polym. Symp.* **1990**, *45*, 289.
- (5) Munk, P.; Du, Q.; Hattam, P. *J. Appl. Polym. Sci., Appl. Polym. Symp.* **1991**, *48*, 293.
- (6) Al-Saigh, Z. Y.; Munk, P. *Macromolecules* **1984**, *17*, 803.
- (7) El-Hibri, M. J.; Munk, P. *Macromolecules* **1988**, *21*, 264.
- (8) Card, T. W.; Al-Saigh, Z. Y.; Munk, P. *J. Chromatogr.* **1984**, *301*, 261.
- (9) Munk, P.; Al-Saigh, Z. Y.; Card, T. W. *Macromolecules* **1985**, *18*, 2196.
- (10) Flory, P. J. *Principles of Polymer Chemistry*; Cornell University Press: Ithaca, NY, 1953.
- (11) Flory, P. J.; Orwoll, R. A.; Vrij, A. J. *J. Am. Chem. Soc.* **1964**, *86*, 3507.
- (12) Sanchez, I. C.; Lacombe, R. H. *Macromolecules* **1978**, *11*, 1145.
- (13) Van Krevelen, D. W. *Properties of Polymers*; Elsevier: Amsterdam, 1972.
- (14) Yaws, C. L. *Thermodynamic and Physical Property Data*; Gulf Publishing Co.: Houston, TX, 1992.
- (15) Reid, R. C.; Prausnitz, J. M.; Poling, B. E. *The Properties of Gases and Liquids*, 4th ed.; McGraw-Hill Book Co.: New York, 1987.
- (16) Smith, B. D.; Srivastava, R. *Thermodynamic Data for Pure Compounds, Part A, Hydrocarbons and Ketones*; Elsevier: New York, 1986.
- (17) Smith, B. D.; Srivastava, R. *Thermodynamic Data for Pure Compounds, Part B, Halogenated Hydrocarbons and Alcohols*; Elsevier: New York, 1986.
- (18) Dreisbach, R. R. *Adv. Chem. Ser.* **1955**, No. 15; **1959**, No. 22; **1961**, No. 29.
- (19) Wilhoit, R. C.; Zwolinski, B. J. *J. Phys. Chem. Ref. Data* **1973**, *2*, Suppl. 1.
- (20) Munk, P.; Card, T. W.; Hattam, P.; El-Hibri, M. J.; Al-Saigh, Z. Y. *Macromolecules* **1987**, *20*, 1278.
- (21) Hattam, P.; Munk, P. *Macromolecules* **1988**, *21*, 2083.
- (22) Shih, H.; Flory, P. J. *Macromolecules* **1968**, *1*, 285.
- (23) Manzini, G.; Crescenzi, V. *Macromolecules* **1975**, *8* (2), 195.
- (24) Takahashi, M.; Shen, M. C.; Taylor, R. B.; Tobolsky, A. V. *J. Appl. Polym. Sci.* **1964**, *8*, 1549.
- (25) Olabisi, O.; Simha, R. *Macromolecules* **1975**, *8*, 206.

Received for review November 17, 1993. Accepted July 25, 1994.[®]

[®] Abstract published in *Advance ACS Abstracts*, September 1, 1994.